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Chemistry 11 KEY
Unit 9 Review

7. Using the Energy Level Diagram for Many Electron Atoms (p. 153 of SW), give the electron configuration for each of the following ions (You may use core notation).

$P^3(44^-)$ [Ne] $3s^2 3p^6$
 $Mg^{2+}(40)$ [Ne] $2s^2 2p^6$
 $Br^-(36)$ [Ar] $4s^2 3d^{10} 4p^6$
 $As^3(34)$ [Ar] $4s^2 3d^{10} 4p^6$
 $Ti^2(38)$ [Ar] $3d^2 4s^0 5p^0$
 $Li^+(7e^-)$ $1s^2$

8. What is the letter which represents the principal quantum number? n
What does it stand for? energy level

9. The elements Ce-Lu are filling up 3 f orbitals.

10. In order to become stable,

an atom of Ca will lose 2 electrons and become the ion Ca^{2+}
an atom of Se will gain 2 electrons and become the ion Se^{2-}
an atom of K will lose 1 electron and become the ion K^+
an atom of Br will gain 1 electron and become the ion Br^-
an atom of N will gain 3 electrons and become the ion N^{3-}
an atom of As will gain 3 electrons and become the ion As^{3-}
an atom of Al will lose 3 electrons and become the ion Al^{3+}
an atom of Te will gain 2 electrons and become the ion Te^{2-}

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